22UCHCT1002

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SEMESTER - I

22UCHCT1002 - Analytical Chemistry

Total Duration : 2 Hrs 30 Mins.

Total Marks : 60

Section B

Answer any **SIX** questions $(6 \times 5 = 30 \text{ Marks})$

- 1. Discuss the first aid techniques used in chemistry laboratories.
- 2. Describe any five-laboratory apparatus used for specific purposes.
- 3. Distinguish between absolute error and relative error.
- 4. a. Outline the principle of iodometric titration.
 - b. Explain the preparation of primary standard solution.
- 5. Discuss the Complex formation titration by EDTA method.
- 6. Explain the following
 - a. DMG b. Cupferon c. Oxime
- 7. Describe the conditions of precipitation.
- 8. Discuss Henderson -Hasselbach equation for buffer solution.

Section C

Answer any **THREE** questions $(3 \times 10 = 30 \text{ Marks})$

- 9. a. List out the precautionary measures to be taken for explosive, carcinogenic and flammable materials.
 - b. How laboratory wastes are treated? Give an example
- 10. a. Explain the method of expressing accuracy.
 - b. Detail the methods of minimising determinate errors.
- 11. a. Discuss the strong acid and strong base by titrimetric method of analysis.
 - b. Explain the following with an example.
 - (i) equivalent weight (ii) normality
- 12. a. Distinguish between Co-precipitate and post precipitation.
 - b. Describe the uses of sequestering agents.
- 13. a. Explain the significances of ka and kb values.
 - b. How the following ions are identified
 - (i) carbonate (ii) sulphate and (iii) cadmium.
